

Unit 10: Acids and Bases

Content Area: **Science**
Course(s): **Chemistry Accelerated**
Time Period: **Marking Period 4**
Length: **1 week**
Status: **Published**

Textbook Resources

Glencoe Science Chemistry Concepts and Applications

Chapter 14: Acids, Bases, and pH

Standards

SCI.9-12.HS-PS1-1	Use the periodic table as a model to predict the relative properties of elements based on the patterns of electrons in the outermost energy level of atoms.
SCI.9-12.HS-PS1-6	Refine the design of a chemical system by specifying a change in conditions that would produce increased amounts of products at equilibrium.
SCI.9-12.HS-PS1-2	Construct and revise an explanation for the outcome of a simple chemical reaction based on the outermost electron states of atoms, trends in the periodic table, and knowledge of the patterns of chemical properties.

Goals/Objectives

- How are acids and bases different?
- Why are acids and bases so important?

Content

- Hydronium and Hydroxide Ions
- Indicators
- Ionization
- Neutralization
- pH

Skills

- • Compare strong v. weak acids and bases
- • Determine pH using various indicators
- • Equate concentration of hydronium to pH

- • Identify, name, and describe the properties of acids and bases
- • Perform a neutralization reaction